

# Electrochemistry

## Question Paper

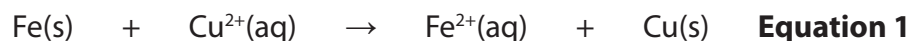
Level	International A Level
Subject	Chemistry
Exam Board	Edexcel
Topic	Chemistry Lab Skills 2
Sub Topic	Electrochemistry
Booklet	Question Paper

**Time Allowed:** 27 minutes  
**Score:** /22  
**Percentage:** /100

**Grade Boundaries:**

A*	A	B	C	D	E	U
>85%	77.5%	70%	62.5%	57.5%	45%	<45%

- 1 A student wishes to measure the  $E_{\text{cell}}$  value of an electrochemical cell in which the following reaction occurs.



The solutions and apparatus available to the student are listed below.

Solution <b>A</b> :	copper(II) sulfate	1.00 mol dm <sup>-3</sup>
Solution <b>B</b> :	iron(II) sulfate	concentration unknown
Solution <b>C</b> :	potassium nitrate	saturated
Solution <b>D</b> :	barium chloride	saturated

Copper foil electrodes  
Iron foil electrodes  
Platinum foil electrodes

Voltmeter **W**: low resistance  
Voltmeter **X**: high resistance  
Ammeter **Y**: low resistance  
Ammeter **Z**: high resistance

Beakers  
Connecting leads  
Crocodile clips  
Strips of filter paper

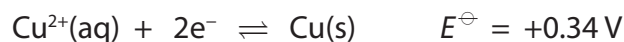
- (a) Draw a labelled diagram of the cell that the student should set up to measure  $E_{\text{cell}}$  for the reaction in **Equation 1**.

Only use items selected from the list above.

(4)

- (b) (i) The student measured  $E_{\text{cell}}$  as +0.79 V. The electrode dipping into the copper(II) sulfate solution was the positive electrode.

For this half-reaction

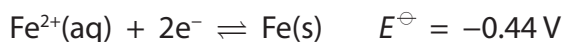


where  $E^{\ominus}$  is the **standard** electrode potential.

Use the above information to calculate the electrode potential ( $E$ ) in the student's cell for the half-reaction



- (ii) For the half-reaction



where  $E^{\ominus}$  is the **standard** electrode potential.

For this half-reaction, the electrode potential ( $E$ ) at a particular concentration is related to the standard electrode potential ( $E^{\ominus}$ ) by the equation

$$E = E^{\ominus} + 0.013 \ln [\text{Fe}^{2+}] \quad \text{Equation 2}$$

where  $\ln$  is the natural logarithm and  $[\text{Fe}^{2+}]$  is the concentration of  $\text{Fe}^{2+}$  ions in  $\text{mol dm}^{-3}$ .

Use **Equation 2**, and your answer to (b)(i), to calculate the concentration of  $\text{Fe}^{2+}$  ions in solution **B**.

(2)

- (c) The concentration of another solution of iron(II) sulfate, **Q**, was found by titration. 25.0 cm<sup>3</sup> samples of **Q** were titrated with a solution of acidified potassium manganate(VII), concentration 0.0300 mol dm<sup>-3</sup>.

The results are as follows:

Titration	Rough	1	2	3
Burette reading (final) / cm <sup>3</sup>	25.00	24.40	24.40	25.70
Burette reading (initial) / cm <sup>3</sup>	1.00	2.10	1.60	3.30
Titre /cm <sup>3</sup>				
Titres used to calculate mean (✓)				

- (i) Complete the table and calculate the mean titre. Indicate with a (✓) the titres that you have used in your calculation.

(2)

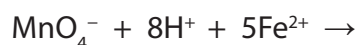
Mean titre .....

- (ii) State the colour change at the end-point.

(1)

- (iii) Complete the equation for the reaction occurring during the titration. State symbols are not required.

(2)



(iv) Calculate the concentration, in  $\text{mol dm}^{-3}$ , of the iron(II) sulfate solution, **Q**.

Give your answer to **three** significant figures.

(4)

(v) The concentration of the iron(II) sulfate solution, **Q**, was also measured on a previous day using the method described in part (a).

The concentration was found to be  $0.157 \text{ mol dm}^{-3}$ .

Calculate the percentage difference between this value and the value you calculated in (c)(iv). You should assume that the correct concentration is  $0.157 \text{ mol dm}^{-3}$ .

(1)

- (vi) In the titration, the volume delivered by the pipette is accurate to  $\pm 0.06 \text{ cm}^3$ .  
Each burette reading is accurate to  $\pm 0.05 \text{ cm}^3$ .

Calculate the percentage error of the pipette for a volume of  $25.00 \text{ cm}^3$  and of the burette for your mean titre.

(2)

Pipette .....

Burette .....

- (vii) Comment on the magnitudes of the values you have calculated in (c)(v) and (c)(vi).

(1)

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- (viii) Suggest why the concentration of iron(II) sulfate in solution **Q** calculated in (c)(iv) is lower than the value given in (c)(v).

(1)

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**(Total for Question 1 = 22 marks)**