

Manufacture and uses includes sulfur dioxide questions

Question Paper 4

Level	IGCSE
Subject	Chemistry
ExamBoard	CIE
Topic	Sulfur
Sub-Topic	Manufacture and uses. Includes sulfur dioxide questions
Paper	(Extended) Theory
Booklet	Question Paper 4

TimeAllowed: 69 minutes

Score: / 57

Percentage: /100

1 Zinc is an important metal. Its uses include making alloys and the construction of dry cells (batteries).

(a) Name an alloy which contains zinc. What is the other metal in this alloy?

name of alloy

other metal in alloy

[2]

(b) The main ore of zinc is zinc blende, ZnS.

(i) The ore is heated in the presence of air to form zinc oxide and sulfur dioxide. Write the equation for this reaction.

..... [2]

(ii) Give a major use of sulfur dioxide.

..... [1]

(c) Zinc can be obtained from zinc oxide in a two step process. Aqueous zinc sulfate is made from zinc oxide and then this solution is electrolysed with inert electrodes. The electrolysis is similar to that of copper(II) sulfate with inert electrodes.

(i) Name the reagent which will react with zinc oxide to form zinc sulfate.

..... [1]

(ii) Complete the following for the electrolysis of aqueous zinc sulfate.

Write the equation for the reaction at the negative electrode.

.....

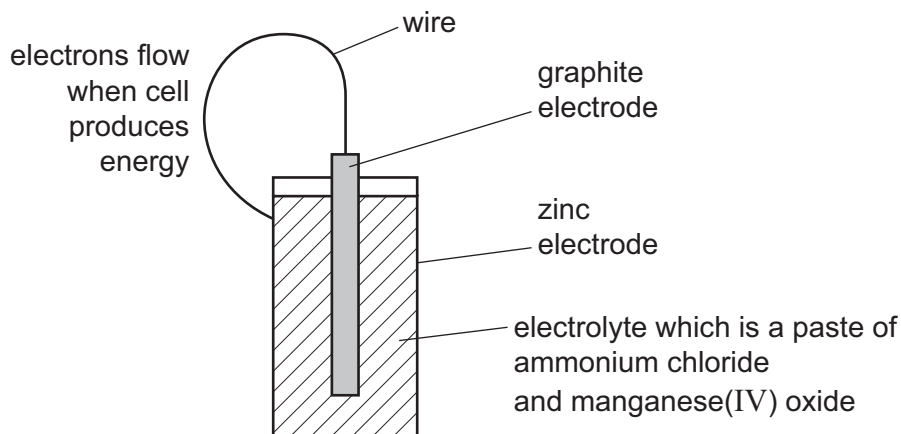
Name the product at the positive electrode.

.....

The electrolyte changes from zinc sulfate to

[3]

- (d) A dry cell (battery) has a central rod, usually made of graphite. This is the positive electrode which is surrounded by the electrolyte, typically a paste of ammonium chloride and manganese(IV) oxide, all of which are in a zinc container which is the negative electrode.



- (i) Draw an arrow on the diagram to indicate the direction of electron flow. [1]

- (ii) Suggest why the electrolyte is a paste.

..... [1]

- (iii) The following changes occur in a dry cell.
For each change, decide if it is oxidation or reduction and give a reason for your choice.

Zn to Zn²⁺

.....

manganese(IV) oxide to manganese(III) oxide

.....

[2]

[Total: 13]

2 Air is a mixture of gases. The main constituents are the elements oxygen and nitrogen.

(a) (i) Name another element in air.

..... [1]

(ii) Give the formula of a compound in unpolluted air.

..... [1]

(b) Common pollutants present in air are the oxides of nitrogen and sulfur dioxide.

(i) How are the oxides of nitrogen formed?

.....
.....
..... [2]

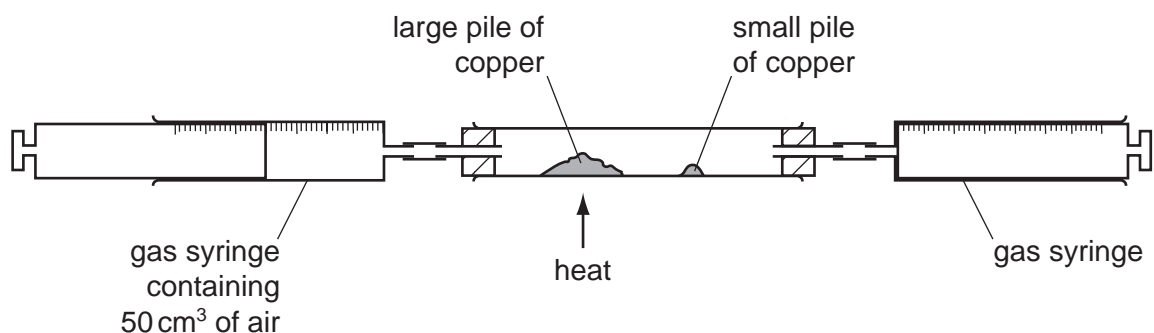
(ii) How is sulfur dioxide formed?

.....
.....
..... [2]

(iii) These oxides are largely responsible for acid rain.
State **two** harmful effects of acid rain.

.....
..... [2]

(c) The percentage of oxygen in air can be determined by the following experiment.



The gas syringe contains 50 cm³ of air. The large pile of copper is heated and the air is passed from one gas syringe to the other over the hot copper. The large pile of copper turns black. The gas is allowed to cool and its volume measured.

The small pile of copper is heated and the remaining gas passed over the hot copper. The copper does not turn black. The final volume of gas left in the apparatus is less than 50 cm³.

(i) Explain why the copper in the large pile turns black.

.....
..... [2]

(ii) Why must the gas be allowed to cool before its volume is measured?

..... [1]

(iii) Explain why the copper in the small pile did not turn black.

..... [1]

(iv) What is the approximate volume of the gas left in the apparatus?

..... [1]

[Total: 13]

3 Zinc is extracted from zinc blende, ZnS.

(a) Zinc blende is heated in air to give zinc oxide and sulphur dioxide. Most of the sulphur dioxide is used to make sulphur trioxide. This is used to manufacture sulphuric acid. Some of the acid is used in the plant, but most of it is used to make fertilisers.

(i) Give another use of sulphur dioxide.

..... [1]

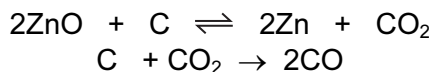
(ii) Describe how sulphur dioxide is converted into sulphur trioxide.

.....
.....
..... [3]

(iii) Name a fertiliser made from sulphuric acid.

..... [1]

(b) Some of the zinc oxide was mixed with an excess of carbon and heated to 1000 °C. Zinc distils out of the furnace.



(i) Name the **two** changes of state involved in the process of distillation.

..... [2]

(ii) Why is it necessary to use an excess of carbon?

.....
..... [2]

(c) The remaining zinc oxide reacts with sulphuric acid to give aqueous zinc sulphate. This is electrolysed with inert electrodes (the electrolysis is the same as that of copper(II) sulphate with inert electrodes).

ions present: $\text{Zn}^{2+}(\text{aq})$ $\text{SO}_4^{2-}(\text{aq})$ $\text{H}^+(\text{aq})$ $\text{OH}^-(\text{aq})$

(i) Zinc forms at the negative electrode (cathode). Write the equation for this reaction.

..... [1]

(ii) Write the equation for the reaction at the positive electrode (anode).

..... [2]

(iii) The electrolyte changes from aqueous zinc sulphate to

..... [1]

(d) Give two uses of zinc.

1.

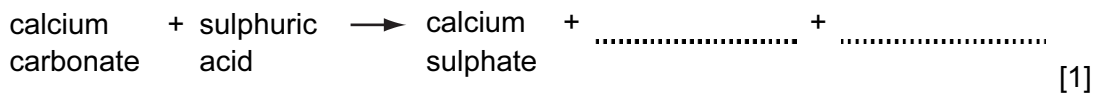
2. [2]

[Total: 15]

4 The Carlsbad caverns in New Mexico are very large underground caves. Although the walls of these caves are coated with gypsum (hydrated calcium sulphate), the caves have been formed in limestone.

(a) It is believed that the caves were formed by sulphuric acid reacting with the limestone.

(i) Complete the word equation.



(ii) Describe how you could test the water entering the cave to show that it contained sulphate ions.

test
result [2]

(iii) How could you show that the water entering the cave has a high concentration of hydrogen ions?

..... [1]

(b) Hydrogen sulphide gas which was escaping from nearby petroleum deposits was being oxidised to sulphuric acid.

(i) Complete the equation for this reaction forming sulphuric acid.



(ii) Explain why all the hydrogen sulphide should be removed from the petroleum before it is used as a fuel.

.....
..... [1]

- (iii) Draw a diagram to show the arrangement of the valency electrons in one molecule of the covalent compound hydrogen sulphide.
Use o to represent an electron from a sulphur atom.
Use x to represent an electron from a hydrogen atom.

[2]

- (c) Sulphuric acid is manufactured by the Contact Process. Sulphur dioxide is oxidised to sulphur trioxide by oxygen.



- (i) Name the catalyst used in this reaction.

..... [1]

- (ii) What temperature is used for this reaction?

..... [1]

- (iii) Describe how sulphur trioxide is changed into sulphuric acid.

.....
..... [2]

- (d) Gypsum is hydrated calcium sulphate, $\text{CaSO}_4 \cdot x\text{H}_2\text{O}$. It contains 20.9% water by mass. Calculate x.

M_r : CaSO_4 , 136; H_2O , 18.

79.1 g of CaSO_4 = moles

20.9 g of H_2O = moles

x = [3]