



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

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**CHEMISTRY**

**0620/11**

Paper 1 Multiple Choice

**May/June 2012**

**45 Minutes**

Additional Materials:      Multiple Choice Answer Sheet  
   Soft clean eraser  
   Soft pencil (type B or HB is recommended)



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**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

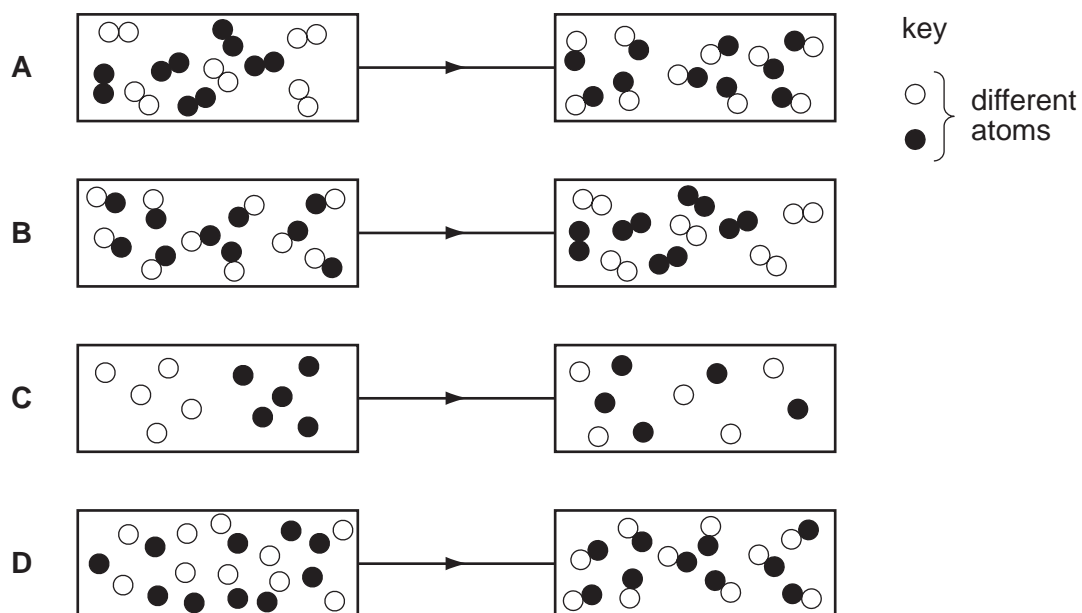
You may use a calculator.

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This document consists of **16** printed pages.



1 Which diagram shows the process of diffusion?



2 Which method is most suitable to obtain zinc carbonate from a suspension of zinc carbonate in water?

- A crystallisation
- B distillation
- C evaporation
- D filtration

3 A student investigates how the concentration of an acid affects the speed of reaction with a 0.5 g mass of magnesium at 30 °C.

The student has a beaker, concentrated acid, water and the apparatus below.

- P a balance
- Q a clock
- R a measuring cylinder
- S a thermometer

Which pieces of apparatus does the student use?

- A P, Q and R only
- B P, Q and S only
- C Q, R and S only
- D P, Q, R and S

- 4 An element Y has the proton number 18.

The next element in the Periodic Table is an element Z.

Which statement is correct?

- A** Element Z has one more electron in its outer shell than element Y.  
**B** Element Z has one more electron shell than element Y.  
**C** Element Z is in the same group of the Periodic Table as element Y.  
**D** Element Z is in the same period of the Periodic Table as element Y.

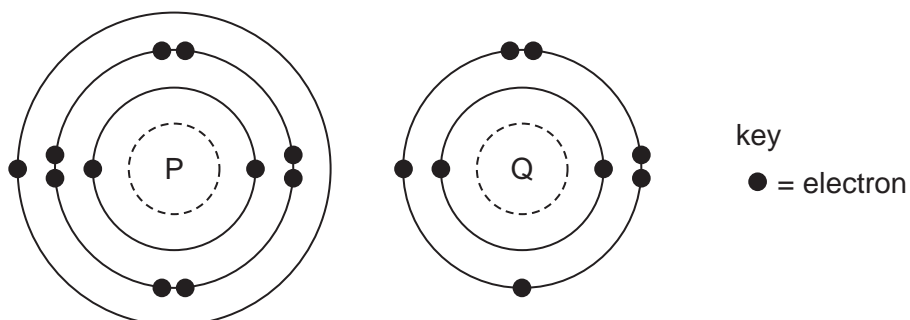
- 5 Which atom has twice as many neutrons as protons?

- A**  ${}^1_1\text{H}$       **B**  ${}^2_1\text{H}$       **C**  ${}^3_1\text{H}$       **D**  ${}^4_2\text{He}$

- 6 Which is a simple covalent molecule?

	conducts electricity		volatile
	when solid	when molten	
<b>A</b>	✓	✓	✗
<b>B</b>	✓	✗	✓
<b>C</b>	✗	✓	✗
<b>D</b>	✗	✗	✓

- 7 The electronic structures of atoms P and Q are shown.

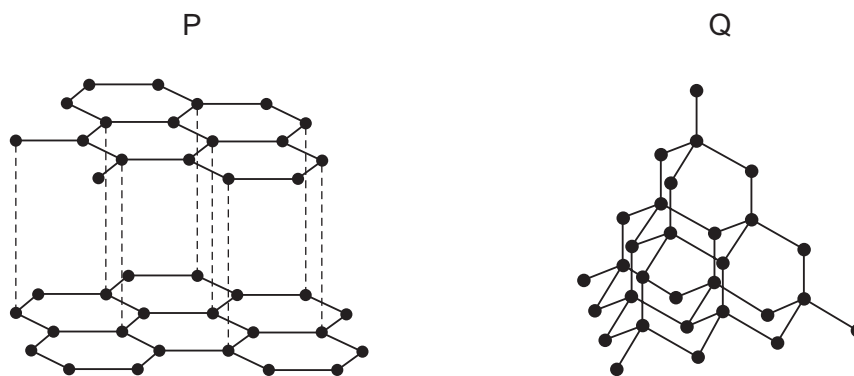


P and Q react to form an ionic compound.

What is the formula of this compound?

- A**  $\text{PQ}_2$       **B**  $\text{P}_2\text{Q}$       **C**  $\text{P}_2\text{Q}_6$       **D**  $\text{P}_6\text{Q}_2$

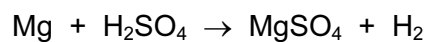
- 8 The diagrams show the structures of two forms, P and Q, of a solid element.



What are suitable uses of P and Q, based on their structures?

	use of solid P	use of solid Q
<b>A</b>	drilling	drilling
<b>B</b>	lubricating	drilling
<b>C</b>	drilling	lubricating
<b>D</b>	lubricating	lubricating

- 9 The equation for the reaction between magnesium and dilute sulfuric acid is shown.



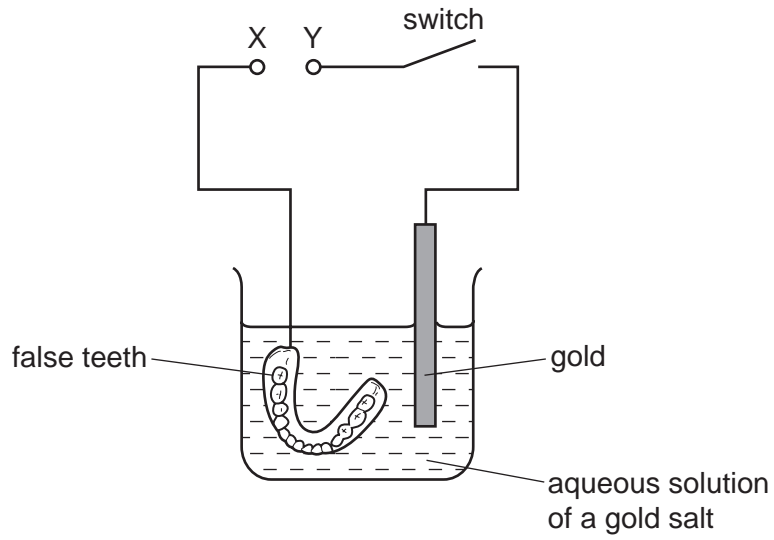
$M_r$  of  $\text{MgSO}_4$  is 120

Which mass of magnesium sulfate will be formed if 12 g of magnesium are reacted with sulfuric acid?

- A** 5g                      **B** 10g                      **C** 60g                      **D** 120g

10 Winston Churchill, a British Prime Minister, had his false teeth electroplated with gold.

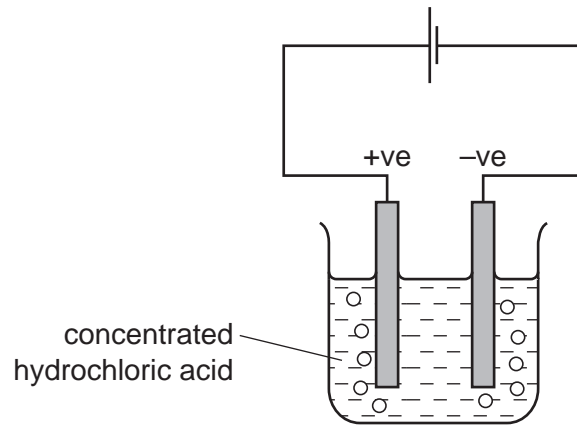
The teeth were coated with a thin layer of carbon and were then placed in the apparatus shown.



Which row is correct?

	terminal X is	the carbon powder could be
<b>A</b>	negative	diamond
<b>B</b>	negative	graphite
<b>C</b>	positive	diamond
<b>D</b>	positive	graphite

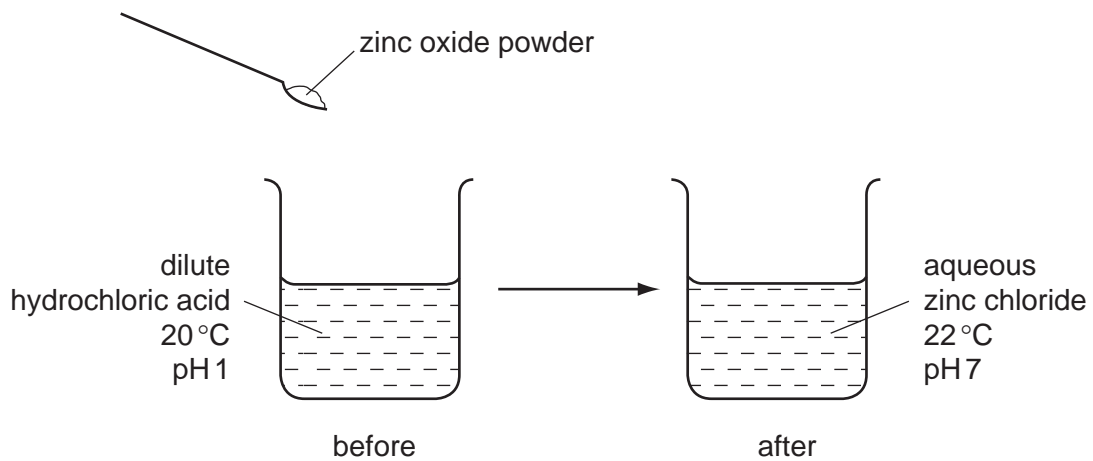
- 11 The diagram shows that two gases are formed when concentrated hydrochloric acid is electrolysed using inert electrodes.



Which row correctly describes the colours of the gases at the electrodes?

	anode (+ve)	cathode (-ve)
<b>A</b>	colourless	colourless
<b>B</b>	colourless	yellow-green
<b>C</b>	yellow-green	colourless
<b>D</b>	yellow-green	yellow-green

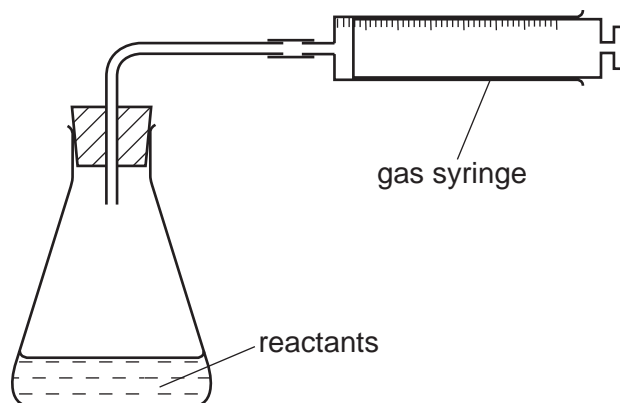
- 12 The diagram shows the reaction between zinc oxide and dilute hydrochloric acid.



Which terms describe the reaction?

	endothermic	neutralisation
<b>A</b>	✓	✓
<b>B</b>	✓	x
<b>C</b>	x	✓
<b>D</b>	x	x

13 The apparatus shown is used to measure the speed of a reaction.



Which equation represents a reaction where the speed can be measured using this apparatus?

- A  $\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
- B  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
- C  $\text{Fe(s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{Cu(s)} + \text{FeSO}_4\text{(aq)}$
- D  $2\text{Na(s)} + \text{Br}_2\text{(l)} \rightarrow 2\text{NaBr(s)}$

14 The element vanadium, V, forms several oxides.

In which change is oxidation taking place?

- A  $\text{VO}_2 \rightarrow \text{V}_2\text{O}_3$
- B  $\text{V}_2\text{O}_5 \rightarrow \text{VO}_2$
- C  $\text{V}_2\text{O}_3 \rightarrow \text{VO}$
- D  $\text{V}_2\text{O}_3 \rightarrow \text{V}_2\text{O}_5$

15 A gas is escaping from a pipe in a chemical plant.

A chemist tests this gas and finds that it is alkaline.

What is this gas?

- A ammonia
- B chlorine
- C hydrogen
- D sulfur dioxide

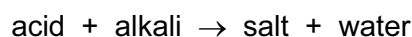
16 The results of three tests on a solution of compound X are shown in the table.

test	result
aqueous sodium hydroxide added	white precipitate formed, soluble in excess
aqueous ammonia added	white precipitate formed, insoluble in excess
acidified silver nitrate added	white precipitate formed

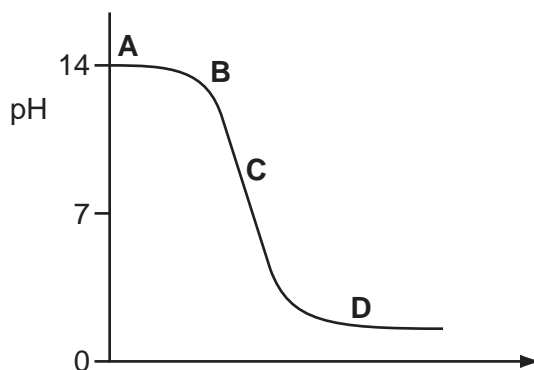
What is compound X?

- A aluminium bromide
- B aluminium chloride
- C zinc bromide
- D zinc chloride

17 The graph shows how the pH changes as an acid is added to an alkali.



Which letter represents the area of the graph where both acid and salt are present?



18 Dilute hydrochloric acid is added to a solid, S.

A flammable gas, G, is formed. Gas G is less dense than air.

What are S and G?

	solid S	gas G
A	copper	hydrogen
B	copper carbonate	carbon dioxide
C	zinc	hydrogen
D	zinc carbonate	carbon dioxide



19 The diagram shows a section of the Periodic Table.

Which element is described below?

'A colourless, unreactive gas that is denser than air.'

			A
	B		
		C	
			D

20 Element X is below iodine in the Periodic Table.

Which row correctly shows the physical state of element X at room temperature and its reactivity compared with that of iodine?

	physical state of element X at room temperature	reactivity compared with that of iodine
A	gas	less reactive
B	solid	less reactive
C	gas	more reactive
D	solid	more reactive

21 Which properties of the element titanium, Ti, can be predicted from its position in the Periodic Table?

	can be used as a catalyst	conducts electricity when solid	has low density	forms coloured compounds
A	✓	✓	x	✓
B	✓	✓	✓	x
C	✓	x	✓	✓
D	x	✓	✓	✓

22 Five elements have proton numbers 10, 12, 14, 16 and 18.

What are the proton numbers of the three elements that form oxides?

- A 10, 12 and 14
- B 10, 14 and 18
- C 12, 14 and 16
- D 14, 16 and 18

23 Which statement about the uses of metals is correct?

- A Aluminium is used in the manufacture of aircraft as it has a high density.
- B Aluminium is used to make food containers as it conducts electricity.
- C Stainless steel for cutlery is made by adding other elements to iron.
- D Stainless steel is used to make chemical reactors as it corrodes readily.

24 Which statement about the extraction of iron from its ore is correct?

- A Iron is more difficult to extract than zinc.
- B Iron is more difficult to extract than copper.
- C Iron is easy to extract because it is a transition metal.
- D Iron cannot be extracted by reduction with carbon.

25 Metal X reacts violently with water.

Metal Y reacts slowly with steam.

Metal Z does not react with dilute hydrochloric acid.

What is the correct order of reactivity of these metals, most reactive first?

- A  $X \rightarrow Y \rightarrow Z$
- B  $X \rightarrow Z \rightarrow Y$
- C  $Z \rightarrow X \rightarrow Y$
- D  $Z \rightarrow Y \rightarrow X$

26 Which property is shown by **all** metals?

- A They are extracted from their ores by heating with carbon.
- B They conduct electricity.
- C They form acidic oxides.
- D They react with hydrochloric acid to form hydrogen.

27 Some uses of water are listed.

- 1 for drinking
- 2 in chemical reactions
- 3 in swimming pools
- 4 in washing

For which uses is it necessary to chlorinate the water?

- A** 1 and 2      **B** 1 and 3      **C** 2 and 4      **D** 3 and 4

28 Coal is a fossil fuel.

Which gas is **not** formed when coal burns?

- A** carbon dioxide  
**B** carbon monoxide  
**C** methane  
**D** sulfur dioxide

29 Which is a use of oxygen?

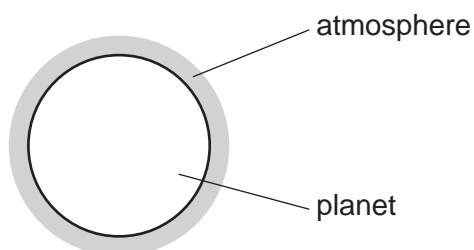
- A** filling balloons  
**B** filling light bulbs  
**C** food preservation  
**D** making steel

30 Fertilisers need to supply crops with three main elements.

Which compound contains all three of these elements?

- A**  $\text{H}_3\text{PO}_4$       **B**  $\text{KNO}_3$       **C**  $\text{NH}_4\text{K}_2\text{PO}_4$       **D**  $\text{NH}_4\text{NO}_3$

31 A new planet has been discovered and its atmosphere has been analysed.



The table shows the composition of the atmosphere.

gas	percentage by volume
carbon dioxide	4
nitrogen	72
oxygen	24

Which gases are present in the atmosphere of the planet in a higher percentage than they are in the Earth's atmosphere?

- A carbon dioxide and oxygen
- B carbon dioxide only
- C nitrogen and oxygen
- D nitrogen only

32 Gas X is a waste gas from digestion in animals.

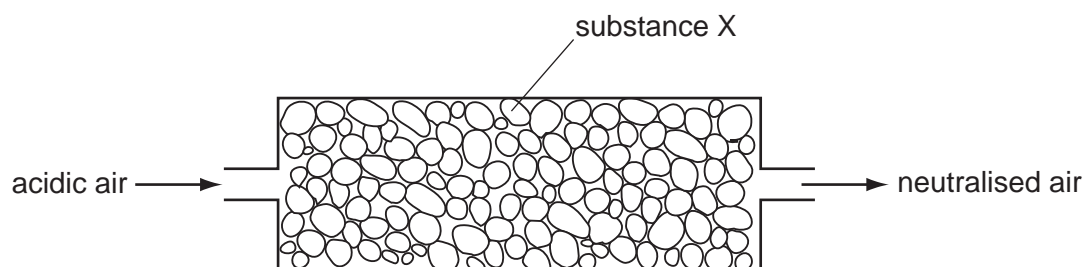
Gas Y is formed when gas X is burnt with a small amount of oxygen.

Gas Z is formed when gas X is burnt with an excess of oxygen.

What are X, Y and Z?

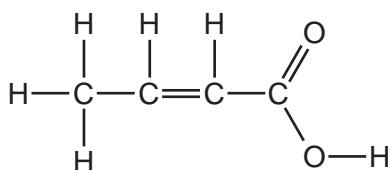
	X	Y	Z
<b>A</b>	carbon dioxide	methane	carbon monoxide
<b>B</b>	carbon monoxide	methane	carbon dioxide
<b>C</b>	methane	carbon dioxide	carbon monoxide
<b>D</b>	methane	carbon monoxide	carbon dioxide

- 33 Air containing an acidic impurity was neutralised by passing it through a column containing substance X.



What is substance X?

- A calcium oxide
  - B sand
  - C sodium chloride
  - D concentrated sulfuric acid
- 34 The structure of a compound is shown.



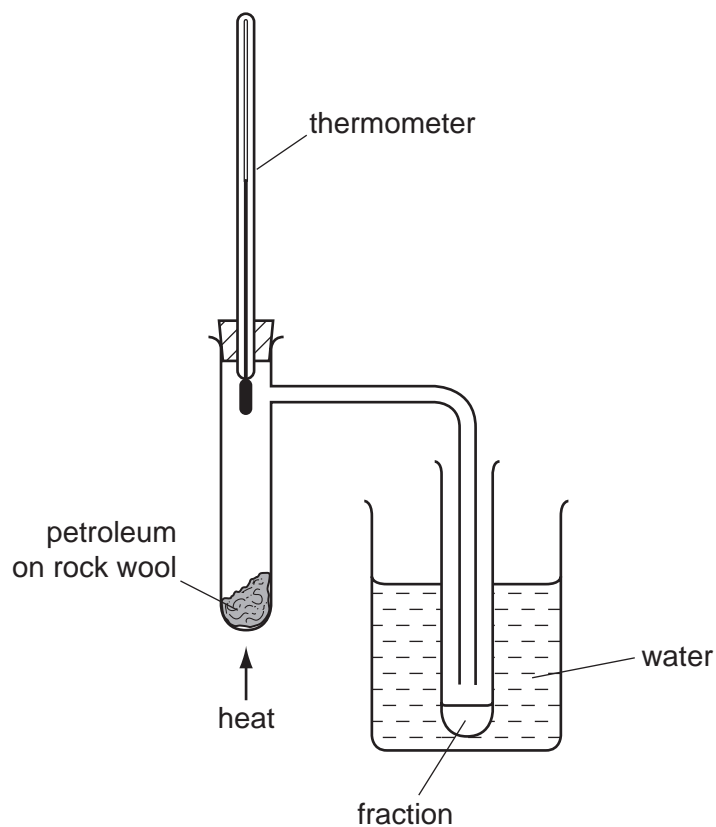
Which functional groups are present in this compound?

	alcohol	alkene	carboxylic acid
<b>A</b>	✓	✓	✓
<b>B</b>	✓	x	x
<b>C</b>	x	✓	✓
<b>D</b>	x	x	✓

- 35 Which fraction from the fractional distillation of petroleum does **not** match its correct use?

	fraction	use
<b>A</b>	fuel oil	domestic heating
<b>B</b>	kerosene	jet fuel
<b>C</b>	naphtha	making roads
<b>D</b>	refinery gas	for heating and cooking

36 The diagram shows apparatus used to separate petroleum into four fractions.



Which fraction contains the smallest hydrocarbon molecules?

fraction	boiling point range / °C
<b>A</b>	up to 70
<b>B</b>	70 to 120
<b>C</b>	120 to 170
<b>D</b>	over 170

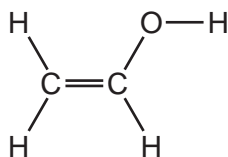
37 When a long chain hydrocarbon is cracked, the following products are produced.

- 1  $C_3H_8$
- 2  $C_2H_4$
- 3  $C_3H_6$
- 4  $C_2H_6$

Which products would decolourise bromine water?

- A** 1 and 4      **B** 2 and 3      **C** 2 only      **D** 3 only

38 PVA is a polymer. The monomer has the structure shown.



To which homologous series does this compound belong?

	alcohols	alkenes
<b>A</b>	✓	✓
<b>B</b>	✓	x
<b>C</b>	x	✓
<b>D</b>	x	x

39 Which equation represents incomplete combustion of ethane?

- A**  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2\text{CO} + 3\text{H}_2$
- B**  $\text{C}_2\text{H}_6 + 2\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2$
- C**  $2\text{C}_2\text{H}_6 + 5\text{O}_2 \rightarrow 4\text{CO} + 6\text{H}_2\text{O}$
- D**  $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$

40 Ethanol is an important chemical produced by the .....1..... of .....2..... .

Which words correctly complete gaps 1 and 2?

	1	2
<b>A</b>	combustion	ethane
<b>B</b>	combustion	glucose
<b>C</b>	fermentation	ethane
<b>D</b>	fermentation	glucose

**DATA SHEET**  
**The Periodic Table of the Elements**

		Group												
		I	II	III	IV	V	VI	VII	VIII	IX	X			
		1 <b>H</b> Hydrogen 1												
7	9													
<b>Li</b> Lithium 3	<b>Be</b> Beryllium 4													
23	24													
<b>Na</b> Sodium 11	<b>Mg</b> Magnesium 12													
39	40													
<b>K</b> Potassium 19	<b>Ca</b> Calcium 20	45	48	51	52	55	56	59	59	64	65			
		<b>Sc</b> Scandium 21	<b>Ti</b> Titanium 22	<b>V</b> Vanadium 23	<b>Cr</b> Chromium 24	<b>Mn</b> Manganese 25	<b>Fe</b> Iron 26	<b>Co</b> Cobalt 27	<b>Ni</b> Nickel 28	<b>Cu</b> Copper 29	<b>Zn</b> Zinc 30			
85	88	89	91	93	96	101	101	103	106	108	112			
<b>Rb</b> Rubidium 37	<b>Sr</b> Strontium 38	<b>Y</b> Yttrium 39	<b>Zr</b> Zirconium 40	<b>Nb</b> Niobium 41	<b>Mo</b> Molybdenum 42	<b>Ru</b> Ruthenium 44	<b>Rh</b> Rhodium 45	<b>Pd</b> Palladium 46	<b>Ag</b> Silver 47	<b>Cd</b> Cadmium 48				
133	137	139	178	181	184	190	190	192	195	197	201			
<b>Cs</b> Caesium 55	<b>Ba</b> Barium 56	<b>La</b> Lanthanum 57	<b>Hf</b> Hafnium 72	<b>Ta</b> Tantalum 73	<b>W</b> Tungsten 74	<b>Os</b> Osmium 76	<b>Ir</b> Iridium 77	<b>Pt</b> Platinum 78	<b>Au</b> Gold 79	<b>Hg</b> Mercury 80				
	226	227												
<b>Fr</b> Francium 87	<b>Ra</b> Radium 88	<b>Ac</b> Actinium 89												
*58-71 Lanthanoid series														
†90-103 Actinoid series														
<table style="width: 100%; border-collapse: collapse;"> <tr> <td style="border: 1px solid black; padding: 2px;">a</td> <td style="border: 1px solid black; padding: 2px;"><b>X</b></td> <td style="border: 1px solid black; padding: 2px;">b</td> </tr> </table> <p>Key a = relative atomic mass X = atomic symbol b = proton (atomic) number</p>												a	<b>X</b>	b
a	<b>X</b>	b												
		140	141	144	150	152	157	159	162	165	167	169	173	175
		<b>Ce</b> Cerium 58	<b>Pr</b> Praseodymium 59	<b>Nd</b> Neodymium 60	<b>Pm</b> Promethium 61	<b>Sm</b> Samarium 62	<b>Gd</b> Gadolinium 64	<b>Tb</b> Terbium 65	<b>Dy</b> Dysprosium 66	<b>Ho</b> Holmium 67	<b>Er</b> Erbium 68	<b>Tm</b> Thulium 69	<b>Yb</b> Ytterbium 70	<b>Lu</b> Lutetium 71
		232	238	238	238	238	238	238	238	238	238	238	238	238
		<b>Th</b> Thorium 90	<b>Pa</b> Protactinium 91	<b>U</b> Uranium 92	<b>Np</b> Neptunium 93	<b>Pu</b> Plutonium 94	<b>Cm</b> Curium 96	<b>Bk</b> Berkelium 97	<b>Cf</b> Californium 98	<b>Es</b> Einsteinium 99	<b>Fm</b> Fermium 100	<b>Md</b> Mendelevium 101	<b>No</b> Nobelium 102	<b>Lr</b> Lawrencium 103
		204	207	209	210	210	210	210	210	210	210	210	210	210
		<b>Tl</b> Thallium 81	<b>Pb</b> Lead 82	<b>Bi</b> Bismuth 83	<b>Po</b> Polonium 84	<b>At</b> Astatine 85	<b>Rn</b> Radon 86	<b>Fr</b> Francium 87	<b>Ra</b> Radium 88	<b>Ac</b> Actinium 89	<b>Th</b> Thorium 90	<b>Pa</b> Protactinium 91	<b>U</b> Uranium 92	<b>Np</b> Neptunium 93

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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