

**June 2004**

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 40**

**SYLLABUS/COMPONENT: 0620/05**

**CHEMISTRY  
Practical**

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	<b>Chemistry – June 2004</b>	<b>0620</b>	<b>5</b>

<b>1</b>	<b>Table of results</b>		
	Experiment 1		
	Temperature boxes completed		1
	Increasing		1
	Comparable to supervisor		1
			<b>[3]</b>
	Experiment 2		
	Temperature boxes completed		1
	Decreasing		1
	Comparable to supervisor		1
			<b>[3]</b>
<b>(a)</b>	All points plotted correctly		4
	(-1 for each incorrect)		
	Smooth line graphs		2
	Labelled		1
			<b>[7]</b>
<b>(b)</b>	<b>(i)</b>	1. Value from graph	1
		2. Value from graph $\pm 0.25$ } No unit only (1)	1
			<b>[2]</b>
	<b>(ii)</b>	1. Exothermic	1
		2. Endothermic	1
			<b>[2]</b>
<b>(c)</b>	Fizz/bubbles/effervescence		1
	Solid disappears		1
			<b>[2]</b>
<b>(d)</b>	Carbonate		1
	Fizz with acid or similar		1
			<b>[2]</b>
<b>(e)</b>	Solid <b>A</b> – value from table/room temperature $\pm 3^{\circ}\text{C}$		1
	Solid <b>B</b> – value from table/room temperature		1
	Reaction finished		1
			<b>[3]</b>
		<b>Sub Total</b>	<b>[24]</b>
<b>2</b>	<b>(a)</b>	White	1
			<b>[1]</b>
	<b>(c)</b>	<b>(i)</b>	White
			Precipitate
			1
			1
			<b>[2]</b>
		Excess – no change	1
			<b>[1]</b>
	<b>(ii)</b>	No precipitate/change	1
			<b>[1]</b>
	<b>(iii)</b>	Paper goes blue	1
		Fizz/bubbles etc	1
		Reference to smell	1
			<b>[3]</b>
	<b>(iv)</b>	pH greater than 7	1
			<b>[1]</b>
	<b>(v)</b>	Milky/cloudy	1
			<b>[1]</b>
<b>(d)</b>	Calcium		1
<b>(e)</b>	Ammonia		1
			<b>[1]</b>

Page 2	Mark Scheme	Syllabus	Paper
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(f)	Limewater	1	
	Carbon dioxide	1	[2]
(g)	Nitrate	1	
	Hydroxide	1	[2]
		<b>Sub Total</b>	<b>[16]</b>
		<b>Total</b>	<b>[40]</b>